Heat Engines, Heat Pumps, and Refrigerators

Getting something useful from heat

Heat *can* be useful

- Normally heat is the end-product of the flow/ transformation of energy
  - remember examples from lecture #5 (coffee mug, automobile, bouncing ball)
  - heat regarded as waste: as useless end result
- Sometimes heat is what we *want*, though
  - hot water, cooking, space heating
- Heat can *also* be coerced into performing "useful" (e.g., mechanical) work
  - this is called a "heat engine"

Heat Engine Concept

- Any time a *temperature difference* exists between two bodies, there is a *potential* for *heat flow*
- Examples:
  - heat flows out of a hot pot of soup
  - heat flows into a cold drink
  - heat flows from the hot sand into your feet
- Rate of heat flow depends on nature of contact and *thermal conductivity* of materials
- If we're clever, we can channel some of this flow of energy into mechanical work

Heat → Work

- We can see examples of heat energy producing other types of energy
  - Air over a hot car roof is lofted, gaining *kinetic energy*
  - That same air also gains *gravitational potential energy*
  - All of our *wind* is driven by temperature differences
  - We already know about *radiative* heat energy transfer
  - Our electricity generation thrives on temperature *differences*: no steam would circulate if everything was at the same temperature

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Power Plant Arrangement

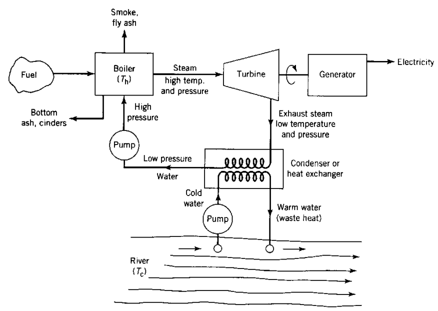

Figure 3.4 A diagram of a fuel-burning electric power plant. Here a river provides cooling water to the condenser, but lake water or a cooling tower could serve the same purpose.

Heat flows from T_h to T_c, turning turbine along the way

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Heat Engine Nomenclature

- The symbols we use to describe the heat engine are:
  - T_h is the temperature of the hot object (typ. in Kelvin)
  - T_c is the temperature of the cold object (typ. in Kelvin)
  - $\Delta T = T_h - T_c$ is the temperature *difference*
  - ΔQ_h is the amount of heat that flows out of the hot body
  - ΔQ_c is the amount of heat flowing into the cold body
  - ΔW is the amount of "useful" mechanical work
  - ΔS_h is the change in *entropy* of the hot body
  - ΔS_c is the change in entropy of the cold body
  - ΔS_{tot} is the total change in entropy (entire system)
  - ΔE is the entire amount of energy involved in the flow

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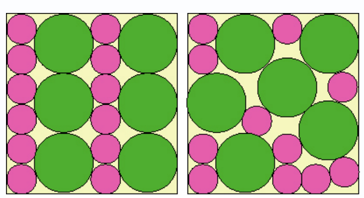
What's this *Entropy* business?

- Entropy is a measure of disorder (and actually quantifiable on an atom-by-atom basis)
  - Ice has low entropy, liquid water has more, steam has a lot



low energy, low T configuration   high energy, high T configuration

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The Laws of Thermodynamics

1. Energy is conserved
2. Total system entropy can never decrease
3. As the temperature goes to zero, the entropy approaches a constant value—this value is zero for a perfect crystal lattice

- The concept of the "total system" is very important: entropy can decrease locally, but it must increase elsewhere by *at least* as much
  - no energy flows into or out of the "total system": if it does, there's more to the system than you thought

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Q

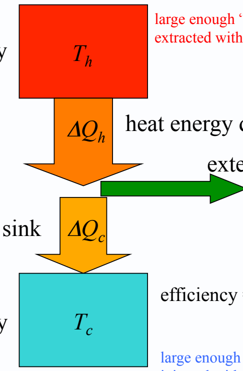
Quantifying heat energy

- We've already seen many examples of quantifying heat
  - 1 kilocalorie is the heat energy associated with raising 1 kg (1 liter) of water 1 ºC
  - In general, $\Delta Q = c_p m \Delta T$, where c_p is the heat capacity
- We need to also point out that a change in heat energy accompanies a change in entropy:

  $\Delta Q = T\Delta S$

  (T expressed in °K, assumed constant)
- Adding heat increases entropy
  - more energy goes into random motions→more randomness (entropy)

How much work can be extracted from heat?



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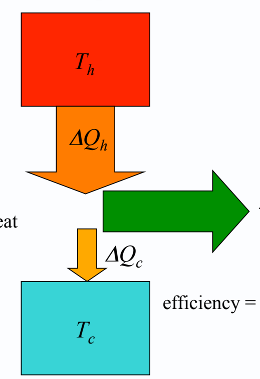
Let's crank up the efficiency



Not so fast…

- The second law of thermodynamics imposes a constraint on this reckless attitude: total entropy must never decrease
- The entropy of the source goes down (heat extracted), and the entropy of the sink goes up (heat added): remember that $\Delta Q = T\Delta S$
  - The gain in entropy in the sink must *at least* balance the loss of entropy in the source

    $\Delta S_{tot} = \Delta S_h + \Delta S_c = -\Delta Q_h/T_h + \Delta Q_c/T_c \geq 0$

    $\Delta Q_c \geq (T_c/T_h)\Delta Q_h$ sets a minimum on ΔQ_c

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What does this entropy limit mean?

- $\Delta W = \Delta Q_h - \Delta Q_c$, so ΔW can only be as big as the minimum ΔQ_c will allow

  $\Delta W_{max} = \Delta Q_h - \Delta Q_{c,min} = \Delta Q_h - \Delta Q_h(T_c/T_h) = \Delta Q_h(1 - T_c/T_h)$
- So the maximum efficiency is:

  maximum efficiency $= \Delta W_{max}/\Delta Q_h = (1 - T_c/T_h) = (T_h - T_c)/T_h$

  this and similar formulas *must* have the temperature in Kelvin
- So perfect efficiency is only possible if T_c is zero (in ºK)
  - In general, this is not true
- As $T_c \rightarrow T_h$, the efficiency drops to zero: no work can be extracted

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Examples of Maximum Efficiency

- A coal fire burning at 825 °K delivers heat energy to a reservoir at 300 °K
  - max efficiency is (825 – 300)/825 = 525/825 = 64%
  - this power station can not possibly achieve a higher efficiency based on these temperatures
- A car engine running at 400 °K delivers heat energy to the ambient 290 °K air
  - max efficiency is (400 – 290)/400 = 110/400 = 27.5%
  - not too far from reality

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Example efficiencies of power plants

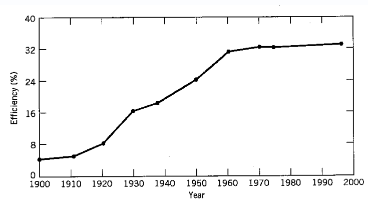

Figure 3.5 Typical efficiency of an electric power plant for converting chemical energy in the fuel into electric energy. The best new plants now achieve nearly 40%. (*Source:* Delbert W. Devins, *Energy: Its Physical Impact on the Environment*, John Wiley and Sons, New York, 1982; and U. S. Energy Information Administration, *Electric Power Annual*, 1996, Volume I.)

Power plants these days (almost all of which are heat-engines) typically get no better than 33% overall efficiency

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What to do with the waste heat (ΔQ_c)?

- One option: use it for space-heating locally

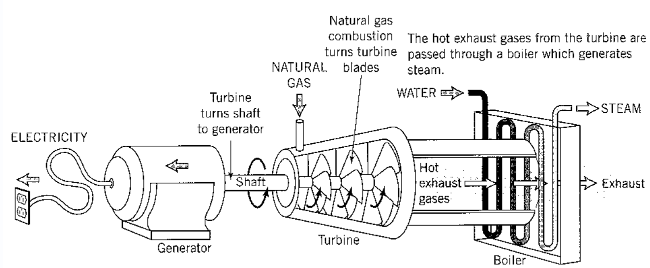

Figure 3.13 A small cogeneration plant that uses the combustion of natural gas to drive a gas turbine coupled to an electric generator. The hot exhaust gases boil water to steam for use in space heating and cooling. (*Source*: Exxon Corporation.)

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Overall efficiency greatly enhanced by cogeneration

Table 3.1 Cogeneration Plant, University of Colorado, Boulder

Fuel	Natural gas
Engine	2 Mitsubishi industrial gas turbines
Generating capacity	32 MW_e
Capital investment	$41,000,000
Construction started	1990
System lifetime	40 to 50 years
Estimated payback time	15 years
Average exported electric power	8 MW_e
Cost of electricity produced	$0.024/kWh
Price of electricity sold	$0.047/kWh
Annual income from electricity sales	$1,600,000
Cost of electricity from public utility	$0.068/kWh
Efficiency for producing electricity	34%
Overall efficiency	70%

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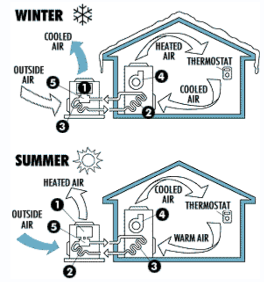
Heat Pumps



Heat Pumps provide a means to very efficiently move heat around, and work both in the winter and the summer

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Heat Pump Diagram

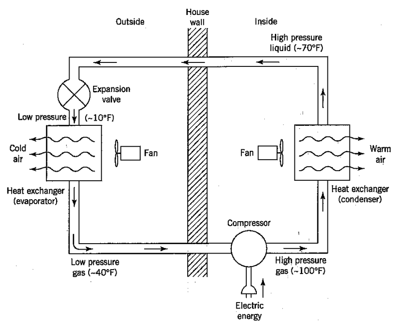

Figure 3.12 An electrically driven heat pump using Freon as a working fluid. In principle, the system becomes an air conditioner if the fluid flow direction is reversed. In practice, the reversal of function is more complex.

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Heat Pumps and Refrigerators: Thermodynamics

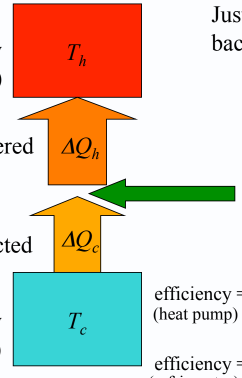
Just a heat engine run backwards…

Hot entity (indoor air)

heat energy delivered

heat energy extracted

Cold entity (outside air or refrigerator)



delivered work:

$\Delta W = \Delta Q_h - \Delta Q_c$

conservation of energy

$$\text{efficiency (heat pump)} = \frac{\Delta Q_h}{\Delta W} = \frac{\text{heat delivered}}{\text{work done}}$$

$$\text{efficiency (refrigerator)} = \frac{\Delta Q_c}{\Delta W} = \frac{\text{heat extracted}}{\text{work done}}$$

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Heat Pump/Refrigerator Efficiencies

- Can work through same sort of logic as before to see that:
  - heat pump efficiency is: $T_h/(T_h - T_c) = T_h/\Delta T$ in ºK
  - refrigerator efficiency is: $T_c/(T_h - T_c) = T_c/\Delta T$ in ºK
- Note that heat pumps and refrigerators are most efficient for small temperature differences
  - hard on heat pumps in very cold climates
  - hard on refrigerators in hot settings

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Example Efficiencies

- A heat pump maintaining 20 ºC when it is –5 ºC outside has a maximum possible efficiency of:

  293/25 = 11.72
  - note that this means you can get almost 12 times the heat energy than you are supplying in the form of work!
  - this factor is called the C.O.P. (coefficient of performance); in practice 2–6 in commercial heat pumps
- A freezer maintaining –5 ºC in a 20 ºC room has a maximum possible efficiency of:

  268/25 = 10.72 → EER = 3.4×10.7 = 36
  - the EER (energy efficiency ratio) is a proxy to this is Btu/hr heat removed per input in Watts
  - 1 Btu/hr is 1055 J/3600 s = 0.293 W → 1 W = 3.4 Btu/hr
  - EER > 3.4 means better than break-even

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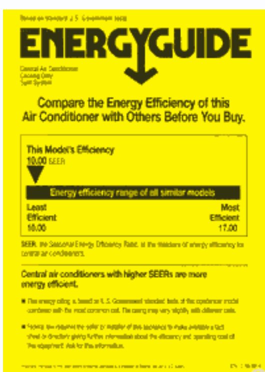
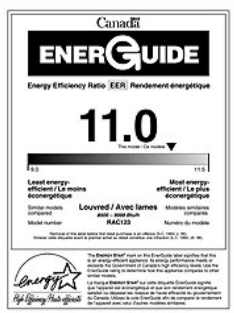
Example Labels (U.S. & Canada)





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Announcements and Assignments

- Chapter 3 goes with this lecture
  - Optional reading on DtM: Heat Pumps Work Miracles
    - http://physics.ucsd.edu/do-the-math/2012/06/heat-pumps-work-miracles/
- HW #3 due Friday 4/26:
  - primarily Chapter 2-related problems: (show work or justify answers!); plus Additional problems (on website)
- HW drop box outside my office (SERF 336) for early turn-in
- Remember that Quizzes happen every week
  - available from Thurs. afternoon until Friday midnight
  - three attempts (numbers change)
    - the better to learn you with

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